

**Daily Tutorial Sheet-8**

**Level - 2**

**96.(B)**  $I.E. = I.E._1 + I.E._2$

**97.(C)**  $\text{moles}(n) = \frac{110}{10^3 \times 7}$ . Energy = moles  $\times$  720 kJ/mole = 11.3 kJ mol<sup>-1</sup>

**98.(D)** Second ionization energy is amount of energy required to take out an electron from the monpositive cation.



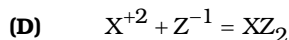
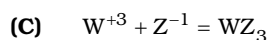
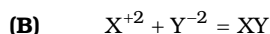
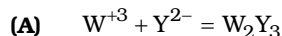
Hence, (5 - 3)

**99. (BC)** W show +3 oxidation state and X show +2 oxidation state.

$Y = ns^2np^4$  Y show -2 oxidation state

$Z = ns^2np^5$ , Z show -1 oxidation state

The compounds that can be formed are :



**100.(B)**  $IE \propto Z_{\text{eff}}$  and EGE of Cl > F, so IE of F<sup>-</sup> < Cl<sup>-</sup>

**101.(B)** Along the period I.E. increases

**102.(B)** Down the group I.E. decreases

**103.(B)** Ist I.E. Mg > Na more  $Z_{\text{eff}}$  & fully filled. But after removing one e<sup>-</sup> from Na, it achieves Noble gas configuration, so removal of dual electron will be more difficult in comparison to 2<sup>nd</sup> e<sup>-</sup> removal from Mg.

**104.(C)** Along the period I.E. increases But I.E of P > S, due to stability of half filled configuration.

**105.(A)**  $I.E. \propto Z_{\text{eff}}$